Name EES : UNIVERSITY OF TLEMCEN **Department: Pharmacy**

SYLLABUS OF MATTER
General pharmaceuticalchemistry

LECTURE TEACHER		**ROUABHI Houria**			
		Studentreception per week			
Email	h_harek@yahoo.fr	Day:	Monday	Hour:	8h30
phone	0541940660	Day:	Thursday	Hour:	8h30
Phone secretary		Day:		Hour:	
Other		Building:		x	

TUTORIALWORKS (Student reception per week)							
NAMES OF TEACHERS	Reception Classroom/office	Session 1		Session 2		Session 3	
		Day	Hour	Day	Hour	Day	Hour
ROUABHI Houria	5	Monday	14h	Monday	14h	x	x
BEMRAH Nawel	6	Wednesday	14h	Wednesday	14h	x	x

PRACTICAL WORKS (Student reception per week)								
NAMES OF TEACHERS	Reception Classroom/office	Session 1			Session 2		Session 3	
		Day	Hour		Day	Hour	Day	Hour
x	x	x	x		x	x	x	x
x	x	x	x		x	x	x	x
x	x	x	x		x	x	x	

COURSE DESCRIPTION	
Goals	General pharmaceutical Chemistry covers the nature ofmatter, stoichiometry, basic chemical reactions, thermodynamic and thermochemistry, atomic structure and theperiodic table, and chemical bonding.
Type of teaching unit	Fundamental
Short course content	The course will focus on chemical structure, bonding and shape as exemplified in the classical model of the atom, Bohr's models, quantum mechanics equation, relation to atomic structure, Hund's and Pauli's exclusion principles: MO and VB approaches to bonding, shapes of atomic and molecular orbitals and Hybridization of atomic orbitals. The Periodic Table, Equilibria in Electrolytes, Acids and

	Bases, Buffers, HandersonHasselberg equation….
Credits of matter	x
Coefficient of matter	3
weighting participation	0%
Weighting diligence	0%
Calculationaverage	(**Exam x 4 + tutorial works)/5**
Skills acquired	At the end of this course the students should be able to: 1. Recognize the state of matter and units of measurements. 2. Recognize Atoms, Molecules, Ions, compounds, Atomic and Electronic Structure and Basic concepts of chemical bonding in addition to Molecular geometry and bondingTheory. 3. Illustrate the Chemical Equilibrium, kinetics, reaction and the factors affecting them. 4. Demonstrate the basic concepts of thermodynamics, solutions and acidity and basicity. 5. Recognize the importance of chemistry to human body. 6. Classify chemical reactions according to their rates and energy. 7. Predict structure, bonding and trends in the behavior of matter using the atomic theory. 8. Differentiate the bonding types, atomic structure, and geometrical shape of molecules. 9. Distinguish between acids and bases and compare the factors affecting the equilibrium and solubility.

BIBLIOGRAPHY	
Books and digital resources	- Understanding General ChemistryBy AtefKorchef - Chimie générale(PATRICK CHAQUIN) - Chimie des solutions (ELISABETH BARDEZ) - Chimie générale(DUNOD) - General Chemistry:Principles and Modern Applications, (Ralph H. Petrucci, William S. Harwood) - Ebbing and Gammon. General Chemistry 11th edition. Belmont: Brooks/Cole CengageLearning, 2016. (Hardcover, Loose Leaf, or eBook), - J.D. Brady, 2000, General Chemistry Principles & Structures, 7th Edition, John Wiley & Sons, New York.
Articles	
Handout.	
WEB Site	https://chemistrynotes.com/pages/entire-year-general-chemistry-notes